Models for the kinetics of enzymes with product inhibition

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Introduction

- Enzymes with competitive inhibition by product
 - Kinetics
 - Mathematical model
- Enzymes with non-competitive inhibition by product
 - Kinetics
 - Mathematical model

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- nature's sustainable catalysts,
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- proteins responsible for thousands of metabolic processes.
- molecular weights: 10,000 to 2,000,000 Dalton
- able to reduce the activation energy of reactions.



Figure: Activation energy. From [2].

Image: A mathematical states and a mathem

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- non-competitive inhibition: the binding of the inhibitor to the enzyme reduces its activity but does not affect the binding of substrate.
- mixed inhibition: the inhibitor can bind to the enzyme at the same time as the substrate.



Figure: Substrate S.

Figure: ES complex.



$$S + E \quad \underbrace{\stackrel{k_1}{\underset{k_{-1}}{\longrightarrow}}} \quad ES \stackrel{k_0}{\longrightarrow} E + P,$$
$$P + E \quad \underbrace{\stackrel{k_2}{\underset{k_{-2}}{\longrightarrow}}} \quad EP,$$

Mathematical model:

$$\begin{aligned} \frac{d[E]}{dt} &= (k_0 + k_{-1})[ES] + k_{-2}[EP] - k_1[E][S] - k_2[E][P], \\ \frac{d[ES]}{dt} &= k_1[E][S] - (k_0 + k_{-1})[ES], \\ \frac{d[EP]}{dt} &= k_2[E][P] - k_{-2}[EP], \\ \frac{d[S]}{dt} &= k_{-1}[ES] + k_1[E][S], \\ \frac{d[P]}{dt} &= k_0[ES] + k_{-2}[EP] - k_2[E][P], \end{aligned}$$

where [A] is the concentration of compound A.

Dimensionless variables:

$$e = \frac{[E]}{e_0}, c_1 = \frac{[ES]}{e_0}, c_2 = \frac{[EP]}{e_0}, s = \frac{[S]}{s_0}, p = \frac{[P]}{s_0}, \tau = e_0k_1t.$$

$$\begin{aligned} \epsilon \frac{dc_1}{d\tau} &= -(s + \hat{k}_0 + \hat{k}_{-1})c_1 - sc_2 + s, \end{aligned} \tag{1} \\ \epsilon \frac{dc_2}{d\tau} &= \hat{k}_2 \left(-pc_1 - (p + \hat{k}_{-2}/\hat{k}_2)c_2 + p \right), \end{aligned} \tag{2} \\ \frac{ds}{d\tau} &= \hat{k}_{-1}c_1 - s(1 - c_1 - c_2)), \end{aligned} \tag{3} \\ \frac{dp}{d\tau} &= \hat{k}_0c_1 + \hat{k}_{-2}c_2 - \hat{k}_2p(1 - c_1 - c_2), \end{aligned} \tag{4}$$

where

$$\epsilon = \frac{e_0}{s_0}, \hat{k}_0 = \frac{k_0}{k_1 s_0}, \hat{k}_{-1} = \frac{k_{-1}}{k_1 s_0}, \hat{k}_2 = \frac{k_2}{k_1}, \hat{k}_{-2} = \frac{k_{-2}}{k_1 s_0}.$$

Image: A math a math

Product formation rate:

$$v = \frac{d[P]}{dt} = \frac{V_{max}[S]}{[S] + K_m \left(1 + \frac{[P]}{K_D}\right)},$$

where

- $V_{max} = k_0 e_0$: maximal rate for enzyme,
- K_m : Michaelis-Menten constant for enzyme,
- K_D : dissociation constant for product.



Figure: Effect of product concentration on the Michaelis-Menten constant for enzyme.





Product formation rate

$$v = rac{d[P]}{dt} = rac{V_{max}}{1 + [P]/K_{D,P}} rac{[S]^2 + A[S]}{[S]^2 + B[S] + C},$$

where

$$\begin{split} & \mathcal{A} = \mathcal{K}_{D,S} + (1+[P]/\mathcal{K}_{D,P})k_{-2}/k_{1}, \\ & \mathcal{B} = \mathcal{K}_{D,S} + \mathcal{K}_{m} + (1+[P]/\mathcal{K}_{D,P})k_{-2}/k_{1}, \\ & \mathcal{C} = \mathcal{K}_{D,S}[\mathcal{K}_{m} + (1+[P]/\mathcal{K}_{D,P})k_{-2}/k_{1}] + k_{0}k_{-2}/k_{1}^{2}, \\ & \mathcal{K}_{D,S} = \frac{k_{-1}}{k_{1}}, \qquad \mathcal{K}_{D,P} = \frac{k_{-2}}{k_{2}}, \qquad \mathcal{K}_{m} = \frac{k_{0}+k_{-1}}{k_{1}}. \end{split}$$

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Figure: Effect of product concentration on V_{max} .

THANK YOU FOR YOUR ATTENTION!

Questions or Comments?!?! ;)

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